ELECTRIC LOAD FACTOR FORMULA Online PDF

electric load factor formula: A Comprehensive Guide to Understanding and Calculating Load Factors

Understanding the concept of load factors is essential for electrical engineers, power plant operators, and energy consumers aiming to optimize energy usage and improve system efficiency. The electric load factor formula provides a quantitative measure of how effectively electrical energy is utilized over a specific period. This article explores the electric load factor formula in depth, explaining its significance, how to calculate it, and its applications in real-world scenarios.

What Is the Electric Load Factor?

The electric load factor is a dimensionless ratio that indicates the efficiency of electrical energy consumption over a given time frame. It reflects the degree to which the maximum demand (peak load) is utilized relative to the total energy consumed.

Definition of Load Factor

The load factor is defined as the ratio of the average load (or demand) over a period to the maximum load during that same period. A higher load factor signifies that the energy is being used more uniformly, leading to better system utilization and reduced operational costs.

Importance of the Load Factor

- Operational Efficiency: Helps utilities optimize generation and distribution.
- Cost Savings: Higher load factors reduce per-unit costs.
- System Planning: Aids in designing infrastructure capable of handling peak loads efficiently.
- Energy Management: Assists consumers in understanding their usage patterns.

Electric Load Factor Formula

The fundamental formula to compute the electric load factor is:

Load Factor (LF) = (Average Load) / (Maximum Load)

Where:

- Average Load is the total energy consumed divided by the period length.
- Maximum Load is the peak demand during that period.

This ratio is often expressed as a decimal or percentage:

- Decimal form: e.g., 0.75
- Percentage form: e.g., 75%

Expressing the Formula in Terms of Energy and Peak Demand

For practical calculations, the formula can be written as:

LF = (Total Energy Consumed in a Period) / (Maximum Demand × Time Period)

Where:

- Total energy is usually measured in kilowatt-hours (kWh).
- Maximum demand is measured in kilowatts (kW).
- Time period is measured in hours (h).

Note: Always ensure units are consistent when applying the formula.

Calculating the Electric Load Factor: Step-by-Step

To calculate the load factor effectively, follow these steps:

- **Determine the Total Energy Consumption**: Obtain the total energy consumed over the period (e.g., daily, monthly, yearly). This data is often recorded by energy meters.
- **Identify the Peak Demand**: Find the maximum load recorded during the period. This can be the highest demand recorded in the energy meter logs or demand readings.
- **Measure the Time Period**: Decide the period over which calculations are to be made (e.g., 24 hours, 30 days).
- **Calculate the Average Load**: Divide the total energy consumed by the total time period:

Average Load = Total Energy Consumed / Total Time Period

- **Apply the Load Factor Formula**: Substitute the values into the load factor formula:

LF = Average Load / Maximum Load

Example Calculation

Suppose a factory consumes 720,000 kWh in a month (30 days), and the maximum demand recorded during that month is 2,400 kW.

- Total energy consumed = 720,000 kWh
- Time period = 30 days × 24 hours/day = 720 hours
- Average load = 720,000 kWh / 720 hours = 1,000 kW

Applying the formula:

LF = 1,000 kW / 2,400 kW = 0.4167 or 41.67%

This indicates that, on average, the factory operates at about 41.67% of its maximum demand, reflecting a moderate load factor.

Factors Influencing the Electric Load Factor

Several factors impact the load factor of a system or consumer:

- **Demand Patterns**: Consistent demand leads to higher load factors.
- **Operational Hours**: Extended operational hours can improve load factor.
- **Type of Consumer**: Industrial users often have higher load factors compared to residential consumers.
- **Energy Efficiency Measures**: Implementation of energy-saving technologies can influence overall load patterns.

Applications of Electric Load Factor

Understanding and calculating the load factor has numerous applications:

1. Utility Load Management

Utilities use load factor data to optimize generation schedules, reduce peak load demands, and plan capacity expansions.

2. Cost Optimization

Consumers with high load factors typically pay lower per-unit charges since they utilize energy more efficiently.

3. Power Plant Design and Operation

Designing power plants with an understanding of expected load factors ensures optimal capacity utilization.

4. Energy Policy and Planning

Government agencies and policymakers analyze load factors to develop sustainable energy strategies.

Limitations and Considerations

While the electric load factor is a valuable metric, it has limitations:

- It does not provide information about the variability of load within the period.
- High load factor does not necessarily mean efficient energy use if total consumption is high.
- It is a historical measure and may not predict future demand patterns accurately.

Therefore, it should be used in conjunction with other metrics like peak demand, load profile, and capacity factor for comprehensive analysis.

Conclusion

The electric load factor formula is a fundamental tool for assessing the efficiency of electrical energy usage over a period. By understanding how to calculate and interpret load factors, utilities, industries, and consumers can make informed decisions to optimize energy consumption, improve system reliability, and reduce costs. Mastery of the load factor concept and its calculation is essential for effective energy management and sustainable power system operations.

Whether you're a professional in the energy sector or a curious consumer, appreciating the significance of the load factor and how it is derived empowers you to better understand the dynamics of electrical systems and their impact on efficiency and cost.

Understanding the Electric Load Factor Formula: A Comprehensive Guide

In the realm of electrical engineering and energy management, the electric load factor formula plays a pivotal role in assessing the efficiency and utilization of electrical power systems. Whether you're an engineer, energy analyst, or facility manager, grasping how to calculate and interpret load factors can lead to more effective energy consumption strategies and cost savings. This article offers an in-depth exploration of the electric load factor formula, its significance, how to compute it, and practical applications.

What Is the Electric Load Factor?

Before diving into the formula, it?s essential to understand the concept of load factor itself. The electric load factor measures the efficiency with which electrical energy is used over a specific period. It compares the average load to the maximum load, providing insight into how consistently electricity is consumed.

Definition:

The electric load factor is a ratio that indicates the portion of the maximum demand that is actually used over a given period, typically a year, month, or day.

Why It Matters:

- It helps utilities and consumers identify opportunities for energy efficiency.
- A higher load factor signifies more consistent and efficient use of power.
- It influences the sizing of generation and distribution infrastructure.

The Electric Load Factor Formula

The electric load factor formula is straightforward but powerful. It mathematically expresses the relationship between the average load and the maximum load during a specific period.

Formula:

Load Factor (LF) = (Average Load) / (Maximum Load)

Expressed as a percentage:

LF (%) = [(Average Load) / (Peak Load)] × 100

Where:

- Average Load is the total energy consumed over the period divided by the number of hours in that period.
- Maximum Load (or peak demand) is the highest level of power demand recorded during the period.

Calculating the Electric Load Factor: Step-by-Step

To accurately compute the electric load factor, follow these steps:

- Determine Total Energy Consumption

- Gather energy consumption data over your chosen period (day, month, year).
- Usually expressed in kilowatt-hours (kWh).
- Example: If a facility consumes 300,000 kWh in a month.

- Calculate the Total Number of Hours

- Decide the period length and note the number of hours in that period.
- For example, a month typically has 30 days × 24 hours = 720 hours.

- Compute the Average Load

- Use the formula:

Average Load (kW) = Total Energy Consumption (kWh) / Total Hours (h)

- Example:

Average Load = 300,000 kWh / 720 h ? 416.67 kW

- Identify the Peak Demand

- Obtain the highest recorded demand during the period, usually from utility billing data or energy monitoring systems.
- Example: The maximum demand recorded is 600 kW.

- Apply the Load Factor Formula

- Plug the numbers into the formula:

LF = (Average Load) / (Maximum Load)

- LF = 416.67 kW / 600 kW ? 0.694 or 69.4%

- Interpret the Result

- A load factor of 69.4% indicates that, on average, the facility uses about 69.4% of its peak demand throughout the period.
- This suggests relatively steady and efficient energy usage.

Practical Applications of the Electric Load Factor

Understanding and calculating the electric load factor has significant implications across various sectors:

Energy Efficiency and Cost Management

- Identify Underutilized Capacity: A low load factor suggests the system is not being used efficiently, prompting measures to improve load management.
- Optimize Infrastructure Sizing: Utilities can better design generation and distribution systems by understanding load patterns.
- Reduce Energy Costs: Facilities can implement demand management strategies during peak periods to lower bills.

Power Generation Planning

- Power plants rely on load factor data to determine the most efficient operation schedules.
- High load factors indicate a consistent demand, enabling more economical operation.

Environmental Impact

- Improved load factors lead to more efficient energy use, reducing unnecessary fuel consumption and emissions.

Factors Influencing the Electric Load Factor

Several factors can affect the load factor of a power system or facility:

- Operational Variability: Seasonal changes, operational hours, and process demands.
- Usage Patterns: Peak hours versus off-peak times.
- Equipment Efficiency: Modern, energy-efficient equipment can help improve load factors.
- Demand Response Initiatives: Programs that encourage shifting usage to off-peak periods.

Strategies to Improve Load Factor

Enhancing the load factor involves increasing the average load relative to the peak demand. Here are some practical strategies:

- Load Shifting: Moving energy-intensive activities to off-peak hours.
- Peak Shaving: Reducing demand during peak periods through demand response.
- Energy Conservation: Promoting efficient usage to maintain steady consumption.
- Process Optimization: Streamlining operations to reduce unnecessary peaks.

Limitations and Considerations

While the electric load factor formula provides valuable insights, it?s important to recognize certain limitations:

- Variability in Data: Inaccurate or incomplete data can skew results.
- Period Selection: The choice of period (daily, monthly, yearly) affects the load factor.
- Peak Demand Measurement: Accurate recording of maximum demand is crucial.
- Contextual Interpretation: A low load factor isn?t always negative; some industries inherently have fluctuating loads.

Conclusion

The electric load factor formula is a fundamental tool in energy management, serving as a benchmark for efficiency and system utilization. By calculating the ratio of average load to maximum load over a defined period, stakeholders can make informed decisions to optimize energy consumption, reduce costs, and enhance sustainability. Whether you're analyzing a manufacturing plant, a commercial building, or a power utility, understanding and applying the load factor formula is essential for effective energy planning and management.

Remember, a higher load factor generally indicates better utilization and efficiency, but context-specific factors should always be considered. Regular monitoring and strategic demand management can lead to significant operational benefits, making the electric load factor a vital metric in the pursuit of smarter energy use.

QuestionAnswer What is the formula for calculating the electric load factor? The electric load factor is calculated using the formula: Load Factor = (Average Load) / (Maximum Load) over a specific period, often expressed as a percentage. Why is the electric load factor important in power system management? The load factor indicates how efficiently electricity is being used over time, helping utilities optimize generation, reduce costs, and improve system reliability. How do you compute the average load in the load factor formula? The average load is computed by dividing the total energy consumed during a period by the total number of hours in that period. Can the electric load factor be greater than 1? No, the load factor cannot be greater than 1 (or 100%), as it represents the ratio of average load to peak load, which cannot exceed the maximum load. What units are used in the electric load factor calculation? The load factor is a dimensionless ratio, often expressed as a percentage, and can be calculated using load values in any consistent units such as kilowatts (kW) and energy in kilowatt-hours (kWh). How can improving the electric load factor benefit utility companies? Improving the load factor leads to better utilization of generation assets, lower operating costs, and more stable grid operations, ultimately reducing electricity costs for consumers.

Related keywords: electric load factor, load factor calculation, electricity consumption, peak load, average load, utility efficiency, power demand, energy usage, load factor formula, electrical engineering

Chemistry I Worksheet Name Calculating Formula Mass

Understanding Chemistry I Worksheet: Calculating Formula Mass

In the foundational course of chemistry, mastering the concept of calculating formula mass is essential for students to comprehend molecular structures, stoichiometry, and chemical reactions. The **chemistry i worksheet name calculating formula mass** serves as an integral educational tool designed to enhance students' understanding of how to determine the mass of chemical formulas accurately. This worksheet not only provides practical exercises but also reinforces theoretical concepts necessary for progressing in chemistry studies.

What Is Formula Mass in Chemistry?

Definition of Formula Mass

Formula mass, also known as molecular weight or molecular mass, refers to the sum of the atomic masses of all atoms in a chemical formula. It is expressed in atomic mass units (amu) or unified atomic mass units (u). Calculating formula mass allows chemists to determine the weight of molecules, which is fundamental for quantitative analysis in chemistry.

Difference Between Formula Mass and Molar Mass

- **Formula Mass:** The sum of atomic masses for the atoms in a chemical formula, expressed per molecule or formula unit.
- **Molar Mass:** The mass in grams of one mole of a substance, numerically equal to the formula mass in amu.

The Importance of Calculating Formula Mass in Chemistry

Calculating the formula mass is crucial for various reasons:

- It helps in determining the molar quantities of substances involved in chemical reactions.
- It aids in converting between mass and number of molecules or moles.
- It is essential for preparing solutions with precise concentrations.
- It supports stoichiometric calculations and reaction yield predictions.

Components of the Chemistry I Worksheet: Calculating Formula Mass

Typical Content and Structure

The worksheet usually contains a variety of exercises designed to build competence in calculating formula mass. These may include:

- Identifying atomic masses of elements from the periodic table.
- Calculating the formula mass of simple molecules and compounds.
- Determining empirical and molecular formulas based on mass data.
- Converting between grams, moles, and number of molecules using formula mass.

Sample Types of Questions

- What is the formula mass of water (H?O)?
- Calculate the formula mass of carbon dioxide (CO?).
- Given a sample containing 12 grams of methane (CH?), find the number of molecules present.
- Determine the empirical formula of a compound with a molar mass of 180 g/mol and composition of 40% carbon, 6.7% hydrogen, and 53.3% oxygen by mass.

Step-by-Step Guide to Calculating Formula Mass

Step 1: Identify the Chemical Formula

Start with the chemical formula of the compound, such as H?O, CO?, or C?H??O?.

Step 2: Find Atomic Masses of Elements

Refer to the periodic table for atomic masses:

- Hydrogen (H): approximately 1.008 amu
- Carbon (C): approximately 12.01 amu
- Oxygen (O): approximately 16.00 amu

Step 3: Multiply Atomic Masses by Subscript Numbers

For each element in the formula, multiply the atomic mass by the number of atoms indicated by the subscript:

- For H?O: (1.008 × 2) + (16.00 × 1) = 2.016 + 16.00 = 18.016 amu

Step 4: Sum All Contributions

Add all the individual atomic masses to obtain the total formula mass:

- H?O: 18.016 amu
- CO?: (12.01 × 1) + (16.00 × 2) = 12.01 + 32.00 = 44.01 amu

Step 5: Express the Result

State the formula mass with appropriate significant figures, considering the precision of atomic weights.

Using the Worksheet to Enhance Learning and Mastery

Practical Tips for Students

- Always double-check the subscript numbers in the chemical formulas.
- Use a periodic table for accurate atomic masses.
- Practice converting between grams, moles, and molecules using the formula mass.
- Learn to interpret word problems and set up calculations systematically.

Benefits of Regular Practice with the Worksheet

- Builds confidence in performing calculations.
- Strengthens understanding of chemical formulas and atomic weights.
- Prepares students for more advanced topics like molar relationships and reaction stoichiometry.
- Encourages critical thinking and problem-solving skills.

Advanced Topics Related to Formula Mass

Empirical and Molecular Formulas

Understanding the difference between empirical and molecular formulas is vital:

- **Empirical Formula:** The simplest whole-number ratio of atoms in a compound.
- **Molecular Formula:** The actual number of atoms in a molecule, which may be a multiple of the empirical formula.

The worksheet may include exercises on deriving empirical and molecular formulas from mass data, involving calculations of molar mass and ratios.

Calculating Molar Mass from Formula Mass

The molar mass in grams per mole is numerically equal to the formula mass in amu, simplifying conversions for laboratory calculations.

Conclusion

The **chemistry i worksheet name calculating formula mass** is an invaluable resource for students learning fundamental chemical concepts. Through structured exercises and step-by-step guidance, students can develop proficiency in calculating the mass of chemical formulas, an essential skill for all branches of chemistry. Regular practice with these worksheets fosters a deeper understanding of atomic structures, molecular compositions, and quantitative analysis, laying the groundwork for success in advanced chemistry topics.

SEO Keywords for Optimization

- Calculating formula mass in chemistry
- Chemistry worksheet for formula mass
- How to find molecular weight
- Atomic masses periodic table
- Stoichiometry and formula mass
- Empirical and molecular formulas
- Chemistry practice exercises
- Understanding chemical formulas
- Mass calculations in chemistry
- Educational resources for chemistry students

Calculating Formula Mass: A Fundamental Skill in Chemistry Education

Understanding the concept of formula mass is essential for students and professionals engaged in the study or practice of chemistry. It forms the backbone of various chemical calculations, from determining molar masses to balancing chemical equations. The "Chemistry I Worksheet: Calculating Formula Mass" serves as a vital educational tool designed to reinforce these foundational skills. This article explores the significance of calculating formula mass, the methods involved, common challenges faced by learners, and strategies to master this critical component of chemical comprehension.

What is Formula Mass? An Introduction

Definition and Significance

The formula mass of a compound is the sum of the atomic masses of all the atoms present in its

chemical formula. It is expressed in atomic mass units (amu), which are numerically equivalent to grams per mole (g/mol). Unlike molecular weight, which is often used for molecules, formula mass applies to any chemical formula, including ionic compounds and complex molecules.

Calculating the formula mass provides a basis for converting between mass and moles, facilitating a deeper understanding of chemical reactions at the molecular level. For example, knowing the formula mass of sodium chloride (NaCl) allows chemists to determine how many grams of salt are needed to obtain a specific number of moles ? a fundamental step in laboratory preparations.

Why is it Important?

The calculation of formula mass underpins many other chemical processes and calculations:

- Mole conversions: Transitioning from grams to moles and vice versa.
- Stoichiometry: Balancing equations and determining reactant/product quantities.
- Percent composition: Analyzing the proportion of each element within a compound.
- Empirical and molecular formulas: Deriving the simplest or actual formulas based on experimental data.

Understanding how to accurately compute formula mass ensures precision in laboratory work and theoretical calculations, making it a core competency for chemistry students.

Fundamentals of Calculating Formula Mass

Atomic Masses: The Building Blocks

The calculation begins with atomic masses, which are standardized values obtained from the periodic table. These are typically expressed in atomic mass units (amu) and represent the average mass of an atom, considering isotopic distribution.

For example:

- Hydrogen (H): approximately 1.008 amu
- Carbon (C): approximately 12.011 amu
- Oxygen (O): approximately 15.999 amu
- Sodium (Na): approximately 22.990 amu
- Chlorine (Cl): approximately 35.45 amu

Having a reliable periodic table or atomic mass reference is crucial for accurate calculations.

Step-by-Step Calculation Approach

Calculating the formula mass involves several systematic steps:

- Identify the elements and their quantities: Determine the number of atoms of each element in the chemical formula.
- Obtain atomic masses: Retrieve the atomic mass for each element.
- Multiply: Multiply the atomic mass of each element by the number of atoms present.
- Sum: Add all the individual totals to obtain the total formula mass.

Example: Calculating the formula mass of glucose (C?H??O?).

- Carbon: 6 atoms × 12.011 amu = 72.066 amu
- Hydrogen: 12 atoms × 1.008 amu = 12.096 amu
- Oxygen: 6 atoms × 15.999 amu = 95.994 amu
- Total formula mass: 72.066 + 12.096 + 95.994 = 180.156 amu

This calculation confirms that glucose has a formula mass of approximately 180.156 amu, which is also its molar mass in grams per mole.

Using Worksheets to Reinforce Learning

Purpose of Chemistry Worksheets

Worksheets like the "Calculating Formula Mass" worksheet serve multiple educational purposes:

- Reinforce theoretical understanding through practical exercises.
- Develop accuracy and speed in calculations.
- Identify common errors and misconceptions.
- Prepare students for standardized testing and laboratory tasks.

These worksheets typically present a series of problems involving various chemical formulas, gradually increasing in complexity to challenge students and solidify their grasp of the concepts.

Features of an Effective Chemistry Worksheet

An effective worksheet should include:

- Clear instructions for each problem.
- A variety of formulas, including simple molecules, ionic compounds, and complex organic structures.
- Step-by-step guided questions to facilitate understanding.
- Space for students to show their calculations.
- Answer keys or solutions to enable self-assessment.

Common Challenges and Troubleshooting

Misinterpretation of Chemical Formulas

One common challenge is misreading chemical formulas, especially when dealing with polyatomic ions, hydrates, or complex organic structures. To address this:

- Students should familiarize themselves with chemical nomenclature.
- Practice parsing formulas, recognizing subscripts, parentheses, and polyatomic groups.
- Use diagrams or models when necessary to visualize the structure.

Inconsistent Atomic Mass Values

Atomic masses can vary slightly depending on the source or the isotopic averages used. To prevent errors:

- Use the same periodic table or atomic mass reference throughout calculations.
- Be aware of the specific values provided in your educational resources.

Calculational Errors

Simple arithmetic mistakes can lead to incorrect results. To mitigate this:

- Double-check each step.
- Use calculator functions for multiplication and addition.
- Practice mental math to improve speed and accuracy.

Advanced Applications and Extensions

Calculating Molar Mass and Moles

Once the formula mass is known, it can be used to find:

- The number of moles in a given mass: moles = mass (g) / molar mass (g/mol)
- The mass of a specific number of moles: mass = moles × molar mass

Percent Composition and Empirical Formulas

Calculating the percent composition involves:

- Dividing the total mass contribution of each element by the formula mass.
- Multiplying by 100 to get percentage.

For example, in glucose:

- % Carbon = (72.066 / 180.156) × 100 ? 40%
- % Hydrogen = (12.096 / 180.156) × 100 ? 6.7%
- % Oxygen = (95.994 / 180.156) × 100 ? 53.3%

These percentages are foundational in deriving empirical formulas, which are the simplest whole-number ratios of elements in a compound.

Educational Strategies to Master Calculating Formula Mass

Hands-On Practice and Repetition

Regular practice using diverse problems helps solidify understanding and improve calculation speed.

Utilizing Visual Aids and Models

Molecular models, diagrams, and periodic table charts can aid in visualizing the composition of compounds.

Peer Collaboration and Group Work

Collaborative problem-solving encourages discussion, clarification of misconceptions, and shared learning experiences.

Incorporating Technology

Educational software and online quizzes can provide instant feedback, reinforcing concepts effectively.

Conclusion: The Importance of Mastering Formula Mass Calculations

Calculating formula mass is more than a routine exercise; it is a gateway to understanding the quantitative aspects of chemistry. The "Chemistry I Worksheet: Calculating Formula Mass" equips students with essential skills that serve as the foundation for advanced chemical concepts and real-world applications. By mastering this skill, students develop a critical analytical mindset, enabling them to approach chemical problems methodically and accurately. As chemistry continues to evolve, the ability to calculate and interpret formula mass remains a cornerstone of scientific literacy, underpinning innovations in research, industry, and education.

In summary, the comprehensive approach to mastering the calculation of formula mass involves understanding atomic weights, systematic procedures, practical application through worksheets, and addressing common challenges. With consistent practice and a solid conceptual foundation, students can confidently navigate the quantitative facets of chemistry, laying the groundwork for success in more complex chemical analyses and research endeavors.

QuestionAnswer What is the purpose of calculating formula mass in chemistry worksheets? Calculating formula mass helps determine the total atomic mass of a compound's formula, which is essential for stoichiometry, balancing equations, and understanding molar relationships. How do you calculate the formula mass of a compound like H?O? To calculate the formula mass of H?O, multiply the atomic mass of hydrogen (1.008 amu) by 2 and add the atomic mass of oxygen (16.00 amu): (2 × 1.008) + 16.00 = 18.016 amu. Why is it important to include the atomic masses of all elements in a chemical formula? Including the atomic masses allows accurate calculation of the compound's formula mass, which is vital for determining molar quantities and converting between mass and moles. What is the difference between molecular mass and formula mass? Molecular mass refers to the mass of a molecule, typically used for covalent compounds, while formula mass is the sum of atomic masses in the empirical formula, often used for ionic compounds or to find molar mass. Can you use periodic table values for calculating formula mass? Yes, you should use the atomic weights from the periodic table to accurately calculate the formula mass of any compound. What steps should I follow to calculate the formula mass of a complex compound? First, identify all elements and their quantities in the formula, then multiply each atomic mass by the number of atoms of that element, and finally sum all these values to get the total formula mass. How does understanding formula mass help in real-world chemistry applications? It helps in preparing precise chemical solutions, calculating yields, and understanding the proportions of reactants and products in chemical reactions. Are there any tools or formulas that simplify calculating formula mass? Yes, using the periodic table for atomic weights and applying straightforward multiplication

and addition makes calculating formula mass quick and accurate. Some calculators or online tools can also assist.

Related keywords: chemistry worksheet, formula mass, molecular weight, atomic mass, molar mass, stoichiometry, chemical formulas, molecular calculations, worksheet exercises, chemistry practice

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